

# Structural Evolution of Reversible Mg Insertion into a Bilayer

## Structure of $V_2O_5 \cdot nH_2O$ Xerogel Material

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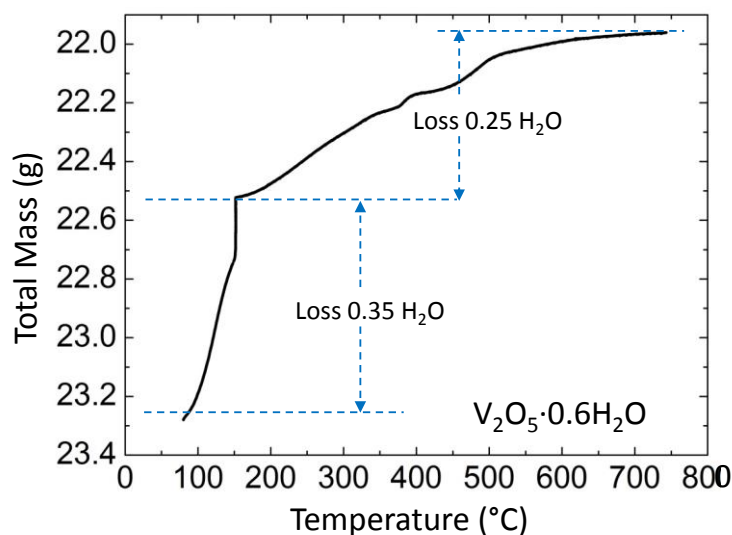
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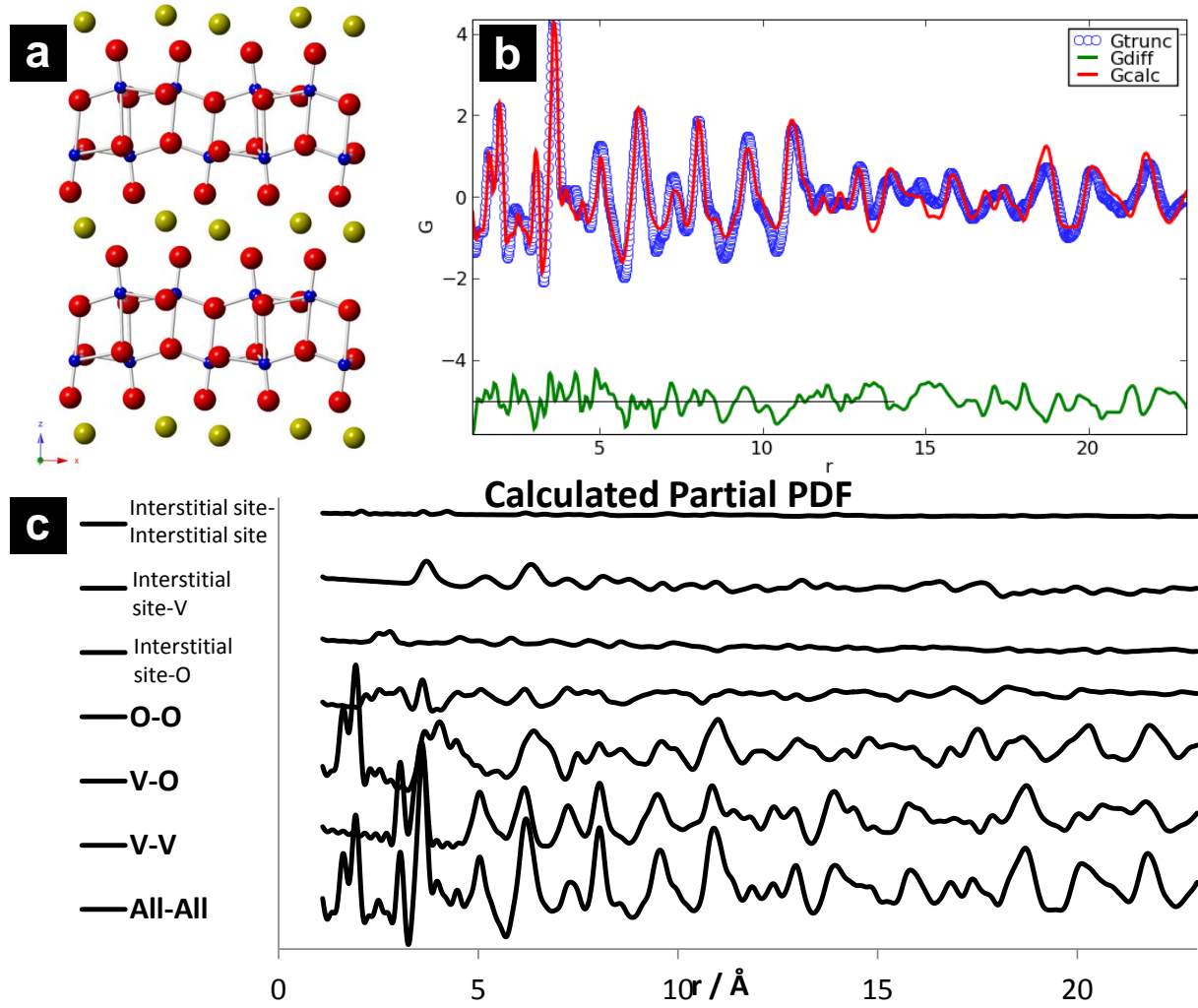
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**KEYWORDS:** Mg battery, bilayer  $V_2O_5$ , Mg intercalation, pair distribution function (PDF), solid state NMR

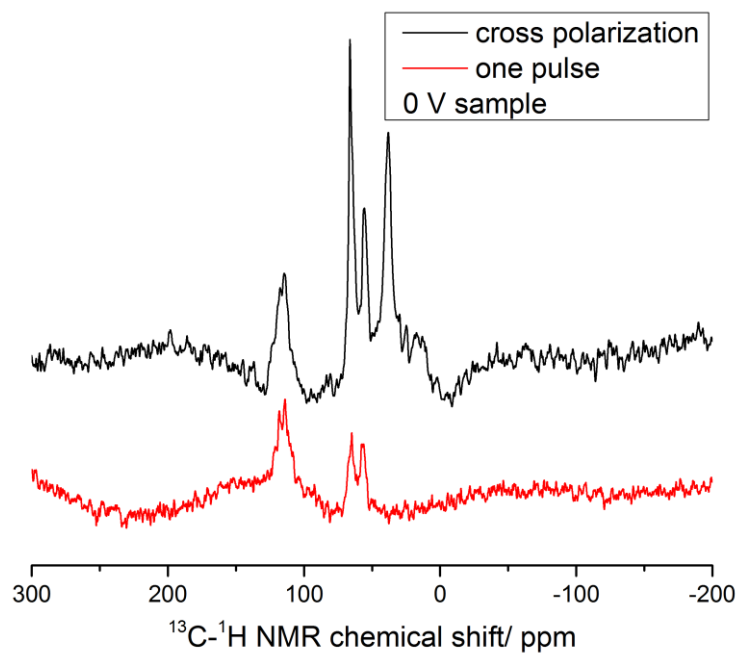
**TGA analysis:** Thermogravimetric analysis (TGA) of the pristine  $V_2O_5 \cdot nH_2O$  xerogel was conducted on a Perkin Elmer Pyris Instrument inside of an argon-filled glove box with  $H_2O$  and  $O_2$  level less than 1 ppm. The weight change profile was characterized by a steep loss program before 80 °C, followed with a gradual weight loss program with 3 °C/min increase of temperature till 800 °C. Figure S1 showed weight loss profile that is composed of two parts, the first slope observed for weight loss before 150 °C which corresponds to weakly bound water and the second slope corresponds to tightly bound water.



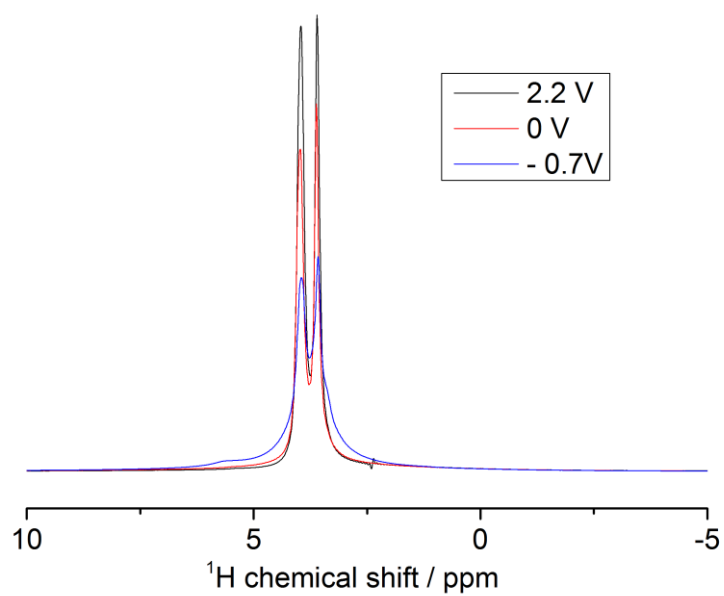
**Figure S1.** Thermogravimetric analysis (TGA) curves for pristine  $V_2O_5 \cdot nH_2O$  xerogel sample. Weight loss corresponding to weakly bond water ( $T < 150$  °C) and tightly bond water ( $T > 150$  °C) is shown.



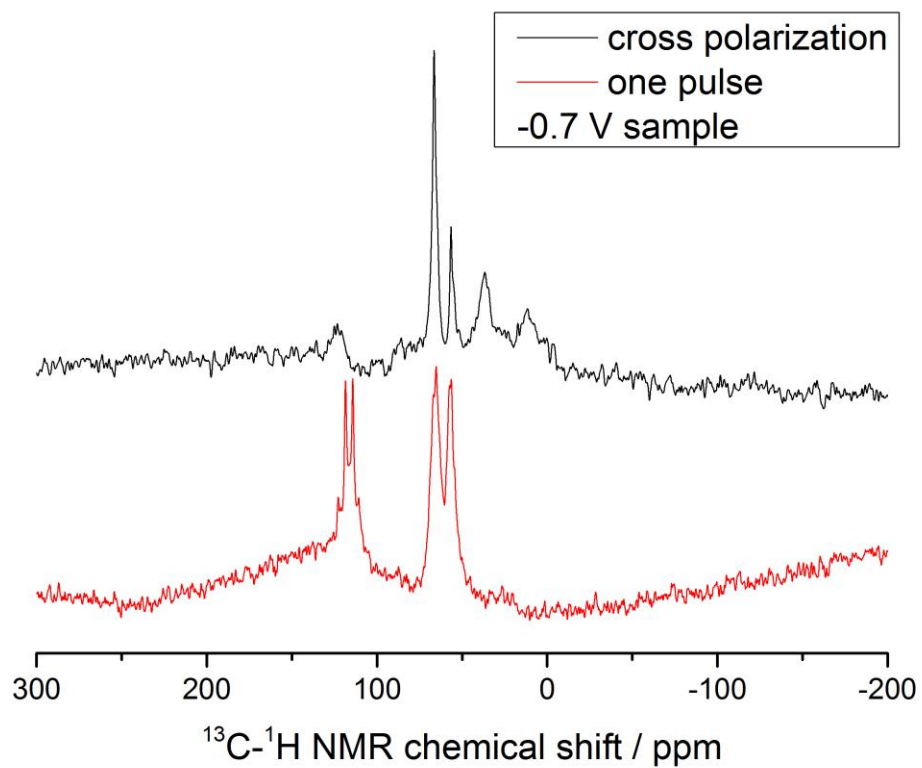
**Figure S2.** Pair distribution function (PDF) analysis for xerogel  $V_2O_5 \cdot nH_2O$ . Structure model is based on  $V_2O_5 \cdot nH_2O$  with interstitial species. **(a)** Double layer structure of xerogel  $V_2O_5 \cdot nH_2O$ , where red, yellow and blue stand for oxygen, water molecules and vanadium, respectively. **(b)** Experimental (hollow blue line) and calculated PDF (solid red line) for  $V_2O_5 \cdot nH_2O$  xerogel. **(c)** Calculated partial PDF for Interstitial site-interstitial site, Interstitial site-V, Interstitial site-O, O-O, V-O, V-V over 20 Å.



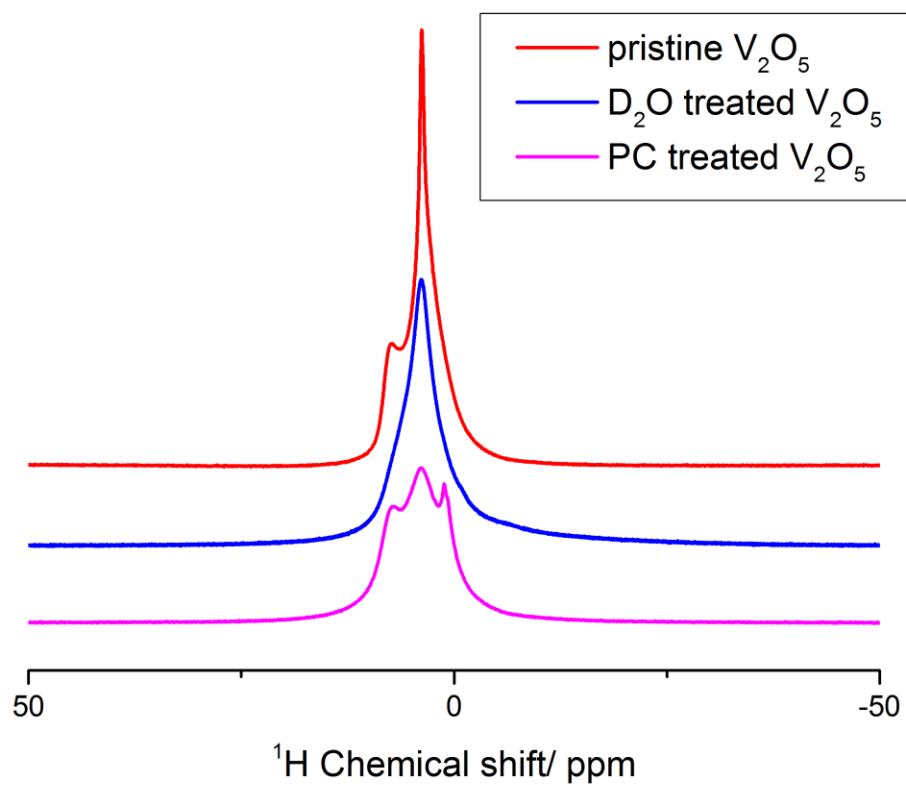
**Figure S3.**  $^{13}\text{C}$  NMR spectra of discharged sample (0 V). black: one pulse experiment, red:  $^1\text{H}$ - $^{13}\text{C}$  cross polarization experiment.



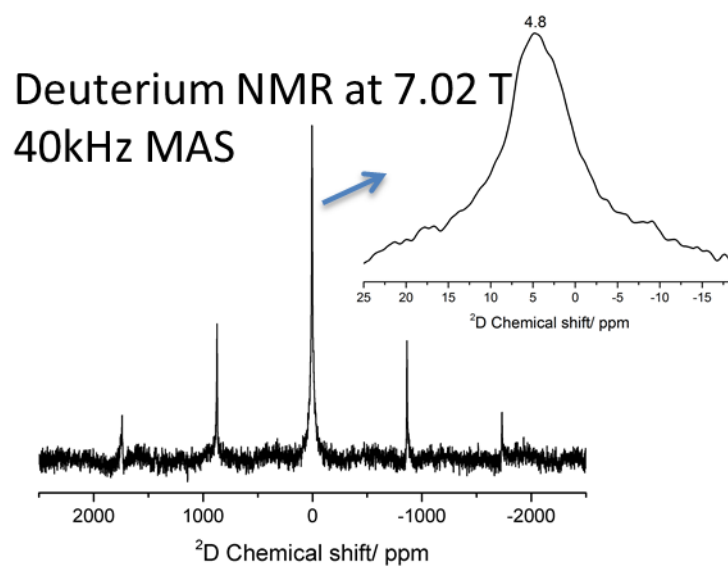
**Figure S4:**  $^1\text{H}$  NMR spectra of a secondary set of charged/discharged sample.



**Figure S5.**  $^{13}\text{C}$  NMR spectra of the -0.7 V sample in Figure S4. Black:  $^1\text{H}$ - $^{13}\text{C}$  cross polarization experiment, red: one pulse experiment.



**Figure S6.**  $^1H$  NMR spectra of pristine  $V_2O_5$ ,  $D_2O$  treated  $V_2O_5$  and PC treated  $V_2O_5$ .



**Figure S7.**  $^2\text{H}$  NMR spectra of  $\text{D}_2\text{O}$  treated  $\text{V}_2\text{O}_5$

**References:**

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- 2 Wang, H.; Senguttuvan, P.; Proffit, D. L.; Pan, B.; Liao, C.; Burrell, A. K.; Vaughey, J. T.; Key, B., Formation of MgO during Chemical magnesiation of Mg-ion battery materials. *ECS Electrochem. Lett.* **2015**, *4*, A90-A93.